

DESCRIBING A SALT PREPARATION

Paper: 1C
Question: 14(e)

Example 1

Your mark

- heat solution on a bunsen burner to evaporate excess solvent. (solute is highly concentrated and left in evaporating dish) leave to cool then
- filter solid, ~~as this will~~ ^{as this will} ~~be~~ ^{form the} crystals
- wash with deionised water (to avoid contamination), and leave to dry in a desiccator and ~~as this~~ hydrate copper (II) sulphate crystals will form.

Example 2

Your mark

They would ~~test~~ place the solution in an evaporating dish and place this on a tripod a gauze, ~~to~~ (with a heat proof mat underneath). They would then use a bunsen burner to heat the ~~solution~~ filtrate (using a roving frame) until half the volume of ~~solution~~ ^{filtrate} evaporates away. They would then leave the filtrate to evaporate and cool at room temperature, ~~leave~~ allowing the crystals to form. ~~to~~ After collecting the crystals, they would use a paper towel to dry any excess filtrate remaining that hasn't crystallised, and then they will be left with dry crystals of hydrated copper (II) sulfate.

Example 3

Your mark

- Once ~~it~~ filtered, the student could gently heat the crystals ~~again~~ using a bunsen burner to evaporate off some of the water
- Then they could use a paper towel to dry any water still on crystals.
- Or they could leave the crystals in a warm and dry place so any remaining water can evaporate off
- This leaves them with a pure, dry sample of the crystals.